

e⁻ (ref)

M_r = 0.00054858 Electron Gas (e⁻)

REFERENCE STATE

0 to 6000 K Ideal Gas

$$\Delta_f H^\circ(0 \text{ K}) = 0 \text{ kJ}\cdot\text{mol}^{-1}$$

$$\Delta_f H^\circ(298.15 \text{ K}) = 0 \text{ kJ}\cdot\text{mol}^{-1}$$

$$S^\circ(298.15 \text{ K}) = 20.979 \pm 0.013 (\pm 2.0) \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$$

Electronic Levels and Quantum Weights	
ϵ , cm ⁻¹	g_i
0	2

Heat of Formation

Zero by definition.

Heat Capacity and Entropy

As shown by Sommerfeld,¹ the electron gas is a degenerate Fermi-Dirac gas and its properties will differ from the classical (Boltzmann) gas. These deviations will increase as the temperature decreases or as the density increases.² Due to the low mass of the electron, these departures from classical behavior will persist to higher temperatures and lower densities than for atomic systems. Under conditions of 1 atm pressure, Gordon³ showed that the deviation of the Fermi-Dirac gas from the Boltzmann gas is negligible above 1250 K. Below this temperature the deviation between classical and quantum statistics will be significant; Mitchell⁴ calculates $S^\circ(298.15 \text{ K}) = 22.72 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ for the Fermi-Dirac gas compared to the classical value of $20.87 \text{ J}\cdot\text{K}^{-1}\cdot\text{mol}^{-1}$ for a pressure of 1 atm.

Despite these known deviations we have chosen to present the classical (Boltzmann) values here since the primary purpose of this table is to serve as a reference state for the calculation of tables of thermodynamic properties for atomic and molecular ions. When ionization equilibria are considered, the density (or partial pressure) of the electron gas will be extremely low except at very high temperatures. At an electron partial pressure of 10⁻⁶ bar, the deviation between classical and quantum statistics will be significant only below 5 K.³ Deviations from classical statistics are encountered at about this same temperature in some atomic and molecular systems at 1 bar pressure, these are usually ignored since the deviations in the enthalpy and entropy are quite small. Therefore, although this ideal-gas table has the formalism of 1 bar as the standard reference state, it should not be applied to real systems where the electron partial pressure exceeds 10⁻⁶ bar. An equivalent statement is that the electron gas will behave as an ideal gas only at electron partial pressures less than about 10⁻⁹ bar, with deviations from ideal gas behavior at higher pressures as a result of the departure from classical statistics.

The thermodynamic functions are calculated here via Boltzmann statistics assuming the electron gas to be an ideal monatomic gas with two equivalent spin states. The relative ionic mass is the electron rest mass as reported in the 1973 CODATA fundamental constants.⁵ The first uncertainty in $S^\circ(298.15 \text{ K})$ represents the uncertainty in the fundamental constants and electron rest mass (i.e., the uncertainty below 10⁻⁶ bar where classical statistics is valid) while the value given in parenthesis represents the uncertainty due to the use of classical rather than quantum statistics at 1 atm pressure. Values below 298.15 K are omitted from the table since the classical calculation (at 1 bar pressure) results in negative values of entropy below 109 K.⁴ The values of the thermodynamic functions agree with those published by Gurvich *et al.*⁶ except for two minor differences. First, the entropy differs by 0.1094 J·K⁻¹·mol⁻¹ because this table uses a standard state pressure of 1 bar, whereas Gurvich *et al.*⁶ used 1 atm. Second, smaller entropy differences occur due to the use of slightly different values of R .

References

- ¹A. Sommerfeld, *Z. Physik* **47**, 1 (1928).
- ²J. E. Mayer and M. G. Mayer, "Statistical Mechanics," Wiley, New York, 495 pp. (1940).
- ³A. R. Gordon, *J. Chem. Phys.* **4**, 678 (1936).
- ⁴A. C. G. Mitchell, *Z. Physik* **50**, 570 (1928).
- ⁵E. R. Cohen and B. N. Taylor, *J. Phys. Chem. Ref. Data* **2**, 663 (1973).
- ⁶L. V. Gurvich, I. V. Veits *et al.*, "Thermodynamic Properties of Individual Substances," 3rd ed., Volume I, Nauka, Moscow, (1978).

T/K	Enthalpy Reference Temperature = T, = 298.15 K		Standard State Pressure = P° = 0.1 MPa		log K _r
	C _p	S° - [G° - H°(T)]/T	H° - H°(T)	Δ _f G°	
0	0	0	0	0	0
100	20.786	20.979	0.038	0.	0.
200	20.786	20.979	1.078	0.	0.
250	20.786	20.979	1.784	0.	0.
300	20.786	20.979	2.117	0.	0.
350	20.786	20.979	2.156	0.	0.
400	20.786	20.979	2.156	0.	0.
450	20.786	20.979	2.156	0.	0.
500	20.786	20.979	2.156	0.	0.
600	20.786	20.979	2.156	0.	0.
700	20.786	20.979	2.156	0.	0.
800	20.786	20.979	2.156	0.	0.
900	20.786	20.979	2.156	0.	0.
1000	20.786	20.979	2.156	0.	0.
1100	20.786	20.979	2.156	0.	0.
1200	20.786	20.979	2.156	0.	0.
1300	20.786	20.979	2.156	0.	0.
1400	20.786	20.979	2.156	0.	0.
1500	20.786	20.979	2.156	0.	0.
1600	20.786	20.979	2.156	0.	0.
1700	20.786	20.979	2.156	0.	0.
1800	20.786	20.979	2.156	0.	0.
1900	20.786	20.979	2.156	0.	0.
2000	20.786	20.979	2.156	0.	0.
2100	20.786	20.979	2.156	0.	0.
2200	20.786	20.979	2.156	0.	0.
2300	20.786	20.979	2.156	0.	0.
2400	20.786	20.979	2.156	0.	0.
2500	20.786	20.979	2.156	0.	0.
2600	20.786	20.979	2.156	0.	0.
2700	20.786	20.979	2.156	0.	0.
2800	20.786	20.979	2.156	0.	0.
2900	20.786	20.979	2.156	0.	0.
3000	20.786	20.979	2.156	0.	0.
3100	20.786	20.979	2.156	0.	0.
3200	20.786	20.979	2.156	0.	0.
3300	20.786	20.979	2.156	0.	0.
3400	20.786	20.979	2.156	0.	0.
3500	20.786	20.979	2.156	0.	0.
3600	20.786	20.979	2.156	0.	0.
3700	20.786	20.979	2.156	0.	0.
3800	20.786	20.979	2.156	0.	0.
3900	20.786	20.979	2.156	0.	0.
4000	20.786	20.979	2.156	0.	0.
4100	20.786	20.979	2.156	0.	0.
4200	20.786	20.979	2.156	0.	0.
4300	20.786	20.979	2.156	0.	0.
4400	20.786	20.979	2.156	0.	0.
4500	20.786	20.979	2.156	0.	0.
4600	20.786	20.979	2.156	0.	0.
4700	20.786	20.979	2.156	0.	0.
4800	20.786	20.979	2.156	0.	0.
4900	20.786	20.979	2.156	0.	0.
5000	20.786	20.979	2.156	0.	0.
5100	20.786	20.979	2.156	0.	0.
5200	20.786	20.979	2.156	0.	0.
5300	20.786	20.979	2.156	0.	0.
5400	20.786	20.979	2.156	0.	0.
5500	20.786	20.979	2.156	0.	0.
5600	20.786	20.979	2.156	0.	0.
5700	20.786	20.979	2.156	0.	0.
5800	20.786	20.979	2.156	0.	0.
5900	20.786	20.979	2.156	0.	0.
6000	20.786	20.979	2.156	0.	0.

PREVIOUS: March 1977 (1 atm)

CURRENT: March 1982 (1 bar)

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